Chemistry

Physical Chemistry

1. Basic principles of chemistry:- Importance of chemistry, Nature of Matter, Properties of Matter and their measurement, Uncertainty in measurements, Laws of chemical combinations, Dalton's Atomic Theory, Atomic and Molecular Masses,Mole concept and molar masses Percentage Composition, Stoichiometry and Stoichiometric Calculations

2. Atomic structure:- Sub atomic Particles, Atomic models, Developments Leading to the Bohr's model of atom, Bohr's Model for hydrogen atom, towards Quantum Mechanical model of the Atom, Quantum mechanical model of Atom, Nature of electromagnetic radiation, photoelectric effect limitations of Bohr's model, Dual nature of matter, de-Broglie's relationship, Heisnberg uncertaintiy principle, various quantum numbers (principal, angular momentum and magnetic quantum numbers) and their significance, shapes of s,p and d – orbitals, electron spin quantum number, Rules for filling electrons in orbitals-aufbau principle , Paulli's exclusion principle and Hund's rule, electronic configuration of elements, extra stability of half- filled and completely filled orbitals.

3. States of Matter:-Intermolecular Forces,Thermal Energy, Intermolecular forces vs thermal interactions, The Gaseous state, The Gas laws, Ideal gas equation, Kinetic Molecular theory of Gases, Liquification of Gases, Liquid state

4. Chemical Bonding and Molecular Structure:-Kossel – Lewis approach to chemical bond formation, concept of ionic and covalent bonds, Ionic Bonding, Formation of Ionic bonds, factors affecting the formation of ionic bonds, calculation of lattice enthalpy, Covalent Bonding, Concept of electronegetavitiy, Fajan's rule, dipole moment, Valence Shell Electron Pair Repulsion(VSEPR) theory and shapes of simple molecules, Quantum mechanical approach to covalent bonding , Valence bond theory – its important features, concept of hybridization involving s,p and d orbitals, Resonance, Molecular Orbital Theory ,LCAOs , types of molecular orbitals (bonding, antibonding), sigma and pi-bonds, molecular orbitals electronic configurations of homonuclear diatomic molecules, concept of bond order, bond length and bond energy , Elementary idea of metallic bonding , Hydrogen bonding and its applications.

5. Basic principles and applications of spectroscopy:- Rotational, vibrational, electronic, Raman, ESR, NMR

6. Thermodynamics:- Fundamental of thermodynamics , System and surroundings, extensive and intensive properties, state functions, types of processes, First law of thermodynamics ,concept of work, heat internal energy and enthalpy , heat capacity, molar heat capacity, Hess's law of constant heat summation , Enthalpies of bond dissociation, combustion ,formation, atomization,sublimation , phase transition, hydration, Ionisationand solution. Second law of thermodynamics, Spontaneity of processes, ΔS of the universe and ΔG of the s system as criteria for spontaneity, ΔG° (standard Gibbs energy change) and equilibrium constant.

7. Equilibrium:- Meaning of equilibrium, concept of dynamic equilibrium. Equilibria involving physical processes: Solid - liquid, liquid - gas and solid - gas equilibria, Henry's

law, general characterics of equilibrium involving physical processes. Equilibria involving chemical process: Law of chemical equilibrium, equilibrium constants (Kp and Kc) and their significance, significance of ΔG and ΔG° in chemical equilibria, factors affecting equilibrium concentration, pressure, temperature, effect of catalyst; Le Chatelier's principle. Ionic equilibrium: Weak and strong electrolytes, ionization of electrolytes, various concepts of acids and bases (Arthenius Bronsted - Lowry and Lewis) and their ionization, acid - base equilibria (including multistage ionization) and ionization constants, ionization of water, pH scale, common ion effect, hydrolysis of salts and pH of their solutions, solubility of sparingly soluble salts and solubility products, buffer solutions.

8. Redox Reactions and Electrochemistry:- Electronic concept of oxidation and reduction, redox reactions, oxidation number, rules for assigning oxidation number balancing of redox reactions. Eectrolytic and metallic conduction, conductance in electrolytic solutions, specific and molar conductivities and their variation with concentration: Kohlrausch's law and its applications. Electrochemical cells - Electrolytic and Galvanic cells, different types of electrodes, electrode potentials including standard electrode potential, half-cell and cell reactions, emf of a Galvanic cell and its measurement; Nernst equation and its applications; Relationship between cell potential and Gibbs' energy change, Dry cell and lead accumulator; Fuel cells.

9. Chemical Kinetics:- Rate of a chemical reactions, factors affecting the rate of reactions: concentration, temperature, pressure and catalyst; elementary and complex reactions, order and molecularity of reactions, rate law, constant and its units, differential and integral forms of zero and first order reactions, their characteristics and half-lives, effect of temperature on rate of reactions - Arrhenius theory, activation energy and its calculation, collision theory of bimolecular gaseous reactions (no derivation).

10. Surface chemistry:- Adsorption - Physisorption and chemisorptions and their characteristics, factors affecting adsorption of gases on solids - Freundlich and Langmuir adsorption isotherms, adsorption from solutions, Colliodal state - distinction among true solutions, colloids and suspensions, classification of colloids - lyophilic, lyophobic; multi molecular, macromolecular and associated colloids (micelles), preparation and properties of colloids - Tyndal effect, Brownian movement, electrophoresis, dialysis, coagulation and flocculation; Emulsions and their characteristics.

11. Solid States-: General Characteristics of solid state, Amorphous and Crystalline Solids, Classification of Crystalline Solids, Crystal Lattices and Unit Cell, Close-Packed Structures, Packing Efficiency, Calculations Involving Unit Cell Dimensions, Imperfections in Solids, Electrical Properties, Magnetic Properties.

12. Concepts of catalysis:- Homogenous and heterogeneous catalysis,

13. Solutions:- Types of Solutions, Expressing concentration of solutions, Solubility, Vapour pressure of liquid solutions, Ideal and Non-ideal solutions, Colligative Properties and Determination of Molar Mass, Abnormal Molar Masses

Inorganic Chemistry

1. Chemical periodicity:- Modern periodic law and present form of the periodic tables, s, p,

d and f block elements, periodic trends in properties of elements atomic and ionic radii, ionization enthalpy, electron gain enthalpy, valence, oxidation states and chemical reactivity.

2. General principles & process of isolation of metals:- Modes of occurrence of elements in nature, minerals, ores; Steps involved in the extraction of metals - concentration, reduction (chemical and electrolytic methods) and refining with special reference to the extraction of Al, Cu, Zn and Fe; Thermodynamic and electrochemical principles involved in the extraction of metals.

3. Hydrogen:- Position of hydrogen in periodic table, isotopes, preparation, properties and uses of hydrogen; Physical and chemical properties of water and heavy water, Structure preparation, reactions and uses of hydrogen peroxide; Hydrogen as a fuel.

4. S-Block elements:- Group-1 and 2 elements introduction, electronic configuration and general trends in physical and chemical properties of elements, anomalous properties of the first element of each group, diagonal relationships. Preparation and properties of some important compounds - sodium carbonate and sodium hydroxide; Industrial uses of lime, limestone Plaster of Paris and cement; Biological significance of Na, K, Mg and Ca.

5. P-Block elements:- Group 13 to Group 18 elements, Electronic configuration general trends in physical and chemical properties of elements across the periods and down the group; unique behavior of the first element in each group. Preparation, properties and uses of boron and aluminium; properties of boric acid, diboron, boron trifluride, aluminium chloride and alums, Allotropes of carbon, catenation; Structure & properties of silicates and zeolites. Properties and uses of nitrogen and phosphorus; Allotrophic forms, structure and uses of ammonia, nitric acid, and PCl3, PCl5; Structures of oxides of phosphorus. Preparation, properties, structures and uses of ozone; Allotropic forms of sulphur, sulphuric acid and structures of oxoacids of sulphur.

6. d-&f Block elements:- Position in periodic table, electronic configurations of d-block elements, general properties of the transition elements (d-Block), some important compounds of transition elements, the lanthanoids, the actinoids, some application of d^3 and f-Block elements. Preparation, properties and uses of $K_2Cr_2O_7$ and $KMnO_4$.

7. Co-ordination compounds & Organometallic compounds:- Introduction to co-ordination compounds, Werner's theory; ligands, co-ordination number, denticity, chelation; IUPAC nomenclature of mononuclear co-ordination compounds, isomerism; Bonding - Valence bond approach and basic ideas of Crystal field theory, colour and magnetic properties; Importance of co-ordination compounds (in qualitative analysis, extraction of metals and in biological systems), Organmetallic comounds-synthesis, bonding and structure, and reactivity. Organometallics in homogenous catalysis. Cages and metal clusters.

8. Environmental Chemistry:- Environmental pollution - Atmospheric, water and soil, Atmospheric pollution-Tropospheric and Stratospheric Tropospheric pollutants- Gaseous pollutants: Oxides of carbon, nitrogen and sulphur, hydrocarbons; their sources, harmful effects and prevention; Green house effect and Global warming; Acid rain; Particulate pollutants: Smoke, dust, smog, fumes, mist; their sources, harmful effects and